

පැරණි නිර්දේශය/பழைய பாடத்திட்டம்/Old Syllabus

ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව
 திணைக்களம் இலங்கைப் பரீட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம்
 Department of Examinations, Sri Lanka Department of Examinations, Sri Lanka Department of Examinations, Sri Lanka Department of Examinations, Sri Lanka Department of Examinations, Sri Lanka
 இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம் இலங்கைப் பரීட்சைத் திணைக்களம்
 Department of Examinations, Sri Lanka

OLD

රසායන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
 கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
 General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව I
 இரசாயனவியல் I
 Chemistry I

02 E I

පැය දෙකයි
 இரண்டு மணித்தியாலம்
 Two hours

Instructions:

- * Periodic Table is provided.
- * This paper consists of 09 pages.
- * Answer all the questions.
- * Use of calculators is not allowed.
- * Write your Index Number in the space provided in the answer sheet.
- * Follow the instructions given on the back of the answer sheet carefully.
- * In each of the questions 1 to 50, pick one of the alternatives from (1), (2), (3), (4), (5) which is correct or most appropriate and mark your response on the answer sheet with a cross (x) in accordance with the instructions given on the back of the answer sheet.

Universal gas constant $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$
 Avogadro constant $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Planck's constant $h = 6.626 \times 10^{-34} \text{ J s}$
 Velocity of light $c = 3 \times 10^8 \text{ ms}^{-1}$

1. Identify the incorrect statement from the following.

- (1) The filling of electrons into orbitals of equal energy is governed by Hund's rule.
- (2) Wave nature of electrons has been shown by diffraction experiments.
- (3) In hydrogen, when electrons fall from high energy levels to the level with principal quantum number, $n = 1$, the line spectrum observed is called Lyman series.
- (4) Atoms absorb or emit radiation in the form of definite small quantities and the smallest quantity is referred to as a photon.
- (5) Two electrons in an orbital must have opposite spins as deduced from Aufbau principle.

2. The number of electrons in the manganese atom (Mn, $Z = 25$) that have quantum numbers $l = 0$ and $m_l = -1$ respectively are,

- (1) 6 and 4 (2) 8 and 12 (3) 8 and 5 (4) 8 and 6 (5) 10 and 5

3. M is an element that belongs to the second period in the Periodic Table. It forms a covalent molecule MCl_3 which has a dipole moment. The group of the Periodic Table to which M belongs is,

- (1) 2 (2) 13 (3) 14 (4) 15 (5) 16

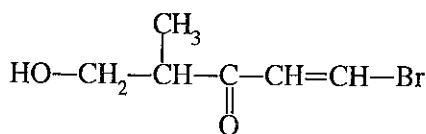
4. The number of unstable Lewis structures that can be drawn for the peroxyntic acid molecule

(formula HNO_4 , $\text{H}-\ddot{\text{O}}-\ddot{\text{O}}-\overset{\text{:O:}}{\parallel}{\text{N}}-\ddot{\text{O}}^-$) is,

- (1) 1 (2) 2 (3) 3 (4) 4 (5) 5

5. The IUPAC name of the given compound is,

- (1) 1-bromo-4-methyl-5-hydroxypent-1-en-3-one
- (2) 5-bromo-1-hydroxy-2-methylpent-4-en-3-one
- (3) 1-bromo-5-hydroxy-4-methylpent-1-en-3-one
- (4) 5-bromo-2-methyl-3-oxopent-4-en-1-ol
- (5) 1-bromo-4-methyl-3-oxopent-1-enol



6. The decreasing order of radii of the species O, O²⁻, F, F⁻, S²⁻, Cl⁻ is,

- (1) S²⁻ > Cl⁻ > O²⁻ > F⁻ > O > F
 (2) S²⁻ > Cl⁻ > O²⁻ > F⁻ > F > O
 (3) Cl⁻ > S²⁻ > O²⁻ > F⁻ > O > F
 (4) Cl⁻ > S²⁻ > F⁻ > O²⁻ > O > F
 (5) S²⁻ > Cl⁻ > O²⁻ > O > F > F

7. A rigid-closed container contains n_1 moles of an ideal gas at temperature T_1 (K) and pressure P_1 (Pa). When an additional amount of the gas was inserted into the container, the new temperature and pressure were T_2 and P_2 , respectively. The total number of moles of the gas now in the container is,

- (1) $\frac{n_1 T_1 P_1}{T_2 P_2}$ (2) $\frac{n_1 T_1 P_1}{T_2 P_1}$ (3) $\frac{T_2 P_2}{n_1 T_1 P_1}$ (4) $\frac{n_1 T_1 P_2}{T_1 P_1}$ (5) $\frac{n_1 T_1 P_1}{T_1 P_2}$

8. The total number of electrons exchanged in the reaction of the oxidation of ethanol (C₂H₅OH) to acetic acid (CH₃COOH) using acidic K₂Cr₂O₇ solution is,

- (1) 6 (2) 8 (3) 10 (4) 12 (5) 14

9. Which compound of the following, can undergo aldol condensation, when reacted with aqueous NaOH?

- (1) CH₃C(=O)OH (2) CH₃C(=O)OCH₃ (3) H-C(=O)OCH₃ (4) CH₃CH₂C(=O)H (5) (CH₃)₃CC(=O)H

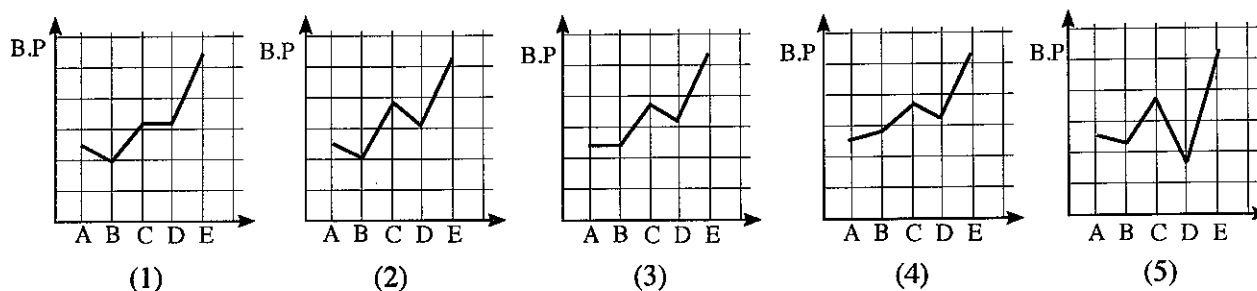
10. AX(s), A₂Y(s) and AZ(s) are sparingly soluble salts in water having K_{sp} values of 1.6×10^{-9} , 3.2×10^{-11} and 9.0×10^{-12} , respectively at 25 °C. Which of the following shows the order of the three saturated solutions of these salts in decreasing concentration of cation A⁺(aq), at 25 °C ?

- (1) AX(s) > A₂Y(s) > AZ(s)
 (2) A₂Y(s) > AX(s) > AZ(s)
 (3) AX(s) > AZ(s) > A₂Y(s)
 (4) A₂Y(s) > AZ(s) > AX(s)
 (5) AZ(s) > A₂Y(s) > AX(s)

11. Consider the following compounds.

	CH ₃ CH ₂ CH ₂ CH ₂ CH ₂ CH ₃	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3\text{CCH}_2\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$	CH ₃ CH ₂ CH ₂ CH ₂ CHO	$\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3\text{CCHO} \\ \\ \text{CH}_3 \end{array}$	CH ₃ CH ₂ CH ₂ CH ₂ CH ₂ OH
	A	B	C	D	E
Relative molecular mass	86	86	86	86	88

Variation of boiling points of these compounds is best shown by,



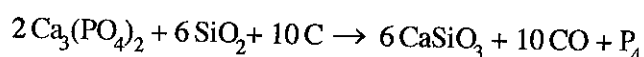
12. The increasing order of covalent character of the chemical species NaCl, Na₂S, KF and KCl is,

- (1) KF < NaCl < KCl < Na₂S
 (2) KCl < NaCl < KF < Na₂S
 (3) KF < KCl < NaCl < Na₂S
 (4) Na₂S < NaCl < KCl < KF
 (5) KF < Na₂S < NaCl < KCl

13. Standard combustion enthalpies of H₂(g), C(s) and CH₃OH(l) at 298 K are -286 kJ mol⁻¹, -393 kJ mol⁻¹ and -726 kJ mol⁻¹, respectively. Enthalpy of vaporization of CH₃OH(l) is +37 kJ mol⁻¹. Enthalpy of formation (kJ mol⁻¹) of one mole of gaseous CH₃OH at 298 K is,

- (1) -276 (2) -239 (3) -202 (4) +84 (5) +202

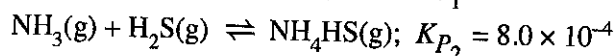
14. Phosphorous can be prepared in an electric furnace as given by the following balanced chemical equation.



When 620 g of Ca₃(PO₄)₂, 180 g of SiO₂ and 96 g of C were reacted, 50 g of P₄ were obtained. Under these conditions, the limiting reagent (reagent that is completely consumed) and percentage yield of P₄ respectively are, (C = 12, O = 16, Si = 28, P = 31, Ca = 40)

- (1) Ca₃(PO₄)₂ and 80.7% (2) SiO₂ and 80.7% (3) C and 50.4%
 (4) SiO₂ and 40.3% (5) C and 25.2%

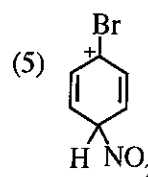
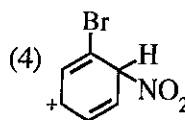
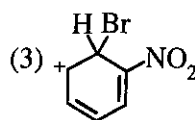
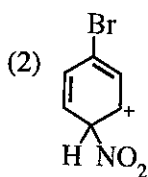
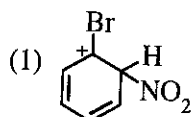
15. Consider the following two equilibria occurring in two separate rigid-closed containers under the same conditions.



Under these conditions K_p for the equilibrium 2H₂S(g) + N₂(g) + 3H₂(g) ⇌ 2NH₄HS(g) is,

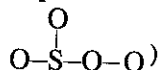
- (1) 5.76 × 10⁻¹² (2) 7.2 × 10⁻¹⁰ (3) 1.92 × 10⁻⁸ (4) 3.40 × 10⁻⁶ (5) 3.75 × 10⁻²

16. Consider the nitration reaction of bromobenzene. Resonance stabilized carbocation intermediates are formed during this reaction. Which of the following is **not** a resonance structure of these intermediates?



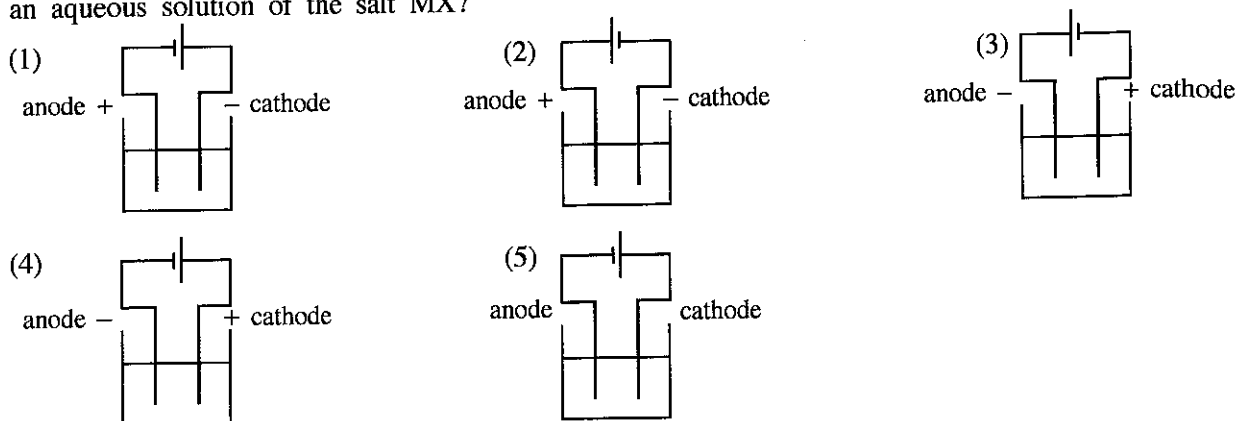
17. A reaction which is non-spontaneous at room temperature and 1 atm pressure becomes spontaneous at high temperature at the same pressure. Which of the following is correct for this reaction at room temperature? (Assume that ΔH and ΔS do not change with temperature and pressure.)

- | | ΔG | ΔH | ΔS |
|-----|----------|----------|----------|
| (1) | Positive | Positive | Positive |
| (2) | Positive | Negative | Negative |
| (3) | Positive | Negative | Positive |
| (4) | Negative | Positive | Negative |
| (5) | Negative | Negative | Negative |

18. The charges on the central sulphur and oxygen atoms in the most acceptable Lewis structure for the SO₄ molecule (skeleton: ) respectively are,

- (1) +1 and zero (2) zero and -1 (3) zero and zero
 (4) +2 and zero (5) +2 and -1

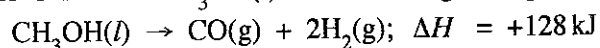
19. Which of the following correctly shows the electrolytic cell constructed for the electrolysis of an aqueous solution of the salt MX?



20. Which of the following statements is correct regarding the reaction between a carboxylic acid and an alcohol to give an ester?

- (1) The overall reaction is a nucleophilic addition reaction of a carbonyl compound.
- (2) It is a reaction in which the alcohol acts as a nucleophile.
- (3) It is a reaction which occurs with the cleavage of the O–H bond of the carboxylic acid.
- (4) It is a reaction which occurs with the cleavage of the C–O bond of the alcohol.
- (5) It is an acid-base reaction.

21. Decomposition of 1 mol of $\text{CH}_3\text{OH}(l)$ occurs at high temperatures as follows.



Which of the following is **incorrect** for the above reaction? (H = 1, C = 12, O = 16)

- (1) The heat absorbed when 1 mol of $\text{CH}_3\text{OH}(g)$ is decomposed is less than 128 kJ.
- (2) Enthalpy of $\text{CO}(g) + 2\text{H}_2(g)$ is higher than the enthalpy of $\text{CH}_3\text{OH}(l)$.
- (3) 128 kJ of heat is released when 1 mol of $\text{CO}(g)$ is formed.
- (4) 128 kJ of heat is absorbed during the decomposition of a mole of reactant.
- (5) 128 kJ of heat is absorbed when 32 g of products are formed.

22. Identify the **incorrect** statement from the following.

- (1) Electron affinity of nitrogen $[\text{N}(g)]$ is positive.
- (2) Dilution of $\text{BiCl}_3(\text{aq})$ solution with water gives a white precipitate.
- (3) H_2S gas can act both as an oxidizing agent and a reducing agent.
- (4) The effective nuclear charge (Z^*) felt by a valence electron in He is less than 2.
- (5) When sulphur reacts with conc. H_2SO_4 , SO_3 gas is produced

23. The concentration of a dilute aqueous solution of a weak acid HA is $C \text{ mol dm}^{-3}$ and its acid dissociation constant is K_a at 298 K. Which of the following expressions gives the pH of the solution at 298 K?

- (1) $\text{pH} = \frac{1}{2} \text{p}K_a - \frac{1}{2} \log C$
- (2) $\text{pH} = -\frac{1}{2} \text{p}K_a - \frac{1}{2} \log C$
- (3) $\text{pH} = -\frac{1}{2} \text{p}K_a + \frac{1}{2} \log C$
- (4) $\text{pH} = -\frac{1}{2} \text{p}K_a - \frac{1}{2} \log (1/C)$
- (5) $\text{pH} = \frac{1}{2} \text{p}K_a - \frac{1}{2} \log (1/C)$

24. The strength of a H_2O_2 solution can be expressed as the volume of O_2 produced at standard temperature and pressure (STP). For example, a litre of 20 **volume strength** H_2O_2 solution will produce 20 litres of O_2 gas at STP ($2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$). (Assume that 1 mole of gas has 22.4 litres volume at STP.)

A bottle labelled X contains H_2O_2 solution. When 25.0 cm^3 of solution X was titrated with 1.0 mol dm^{-3} KMnO_4 in the presence of dilute H_2SO_4 the volume required to reach the end point was 25.0 cm^3 . The volume strength of solution X is,

- (1) 15 (2) 20 (3) 25 (4) 28 (5) 30

25. $\text{M}(\text{OH})_2(\text{s})$ is a sparingly water soluble salt formed by the reaction between $\text{M}^{2+}(\text{aq})$ and $\text{OH}^-(\text{aq})$ ions at 298 K. The solubility (mol dm^{-3}) of $\text{M}(\text{OH})_2(\text{s})$ in water at $\text{pH} = 5$ is, ($K_{sp}\text{M}(\text{OH})_2 = 4.0 \times 10^{-36}$ at 298 K).

- (1) $\sqrt{2} \times 10^{-18}$ (2) 2×10^{-18} (3) 1×10^{-18} (4) $\sqrt[3]{2} \times 10^{-12}$ (5) 1×10^{-12}

26. Which of the following correctly denotes the standard galvanic cell constructed by using a standard hydrogen electrode, a standard Mg-electrode and a salt-bridge at 298 K?

- (1) $\text{Mg}(\text{s}) \mid \text{Mg}^{2+}(\text{aq}, 1.00 \text{ mol dm}^{-3}) \parallel \text{H}^+(\text{aq}, 1.00 \text{ mol dm}^{-3}) \mid \text{H}_2(\text{g}) \mid \text{Pt}(\text{s})$
 (2) $\text{Pt}(\text{s}) \mid \text{H}_2(\text{g}) \mid \text{H}^+(\text{aq}, 1.00 \text{ mol dm}^{-3}) \parallel \text{Mg}^{2+}(\text{aq}, 1.00 \text{ mol dm}^{-3}) \mid \text{Mg}(\text{s})$
 (3) $\text{Mg}(\text{s}), \text{Mg}^{2+}(\text{aq}, 1.00 \text{ mol dm}^{-3}) \parallel \text{H}^+(\text{aq}, 1.00 \text{ mol dm}^{-3}) \mid \text{H}_2(\text{g}) \mid \text{Pt}(\text{s})$
 (4) $\text{Mg}(\text{s}) \mid \text{Mg}^{2+}(\text{aq}, 1.00 \text{ mol dm}^{-3}), \text{H}^+(\text{aq}, 1.00 \text{ mol dm}^{-3}), \text{H}_2(\text{g}) \mid \text{Pt}(\text{s})$
 (5) $\text{Pt}(\text{s}), \text{H}_2(\text{g}) \mid \text{H}^+(\text{aq}, 1.00 \text{ mol dm}^{-3}) \parallel \text{Mg}^{2+}(\text{aq}, 1.00 \text{ mol dm}^{-3}), \text{Mg}(\text{s})$

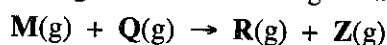
27. The following procedure was carried out at 298 K to determine the distribution coefficient K_D of a monobasic organic acid between dichloromethane and water. 50.00 cm^3 of a 0.20 mol dm^{-3} aqueous solution of acid were mixed vigorously with 10.00 cm^3 of dichloromethane and the two layers were allowed to separate. Thereafter, the dichloromethane layer in the bottom of the flask was drained out. 10.00 cm^3 of 0.02 mol dm^{-3} $\text{NaOH}(\text{aq})$ solution were required to neutralize the acid remaining in the aqueous layer. (Assume that the acid does not dimerize in the organic phase.) K_D of the acid between **dichloromethane and water** at 298 K is,

- (1) 0.05 (2) 0.25 (3) 4.00 (4) 20.00 (5) 245.00

28. A reaction $\text{C}_2\text{H}_4(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ occurs in a rigid-closed container at a given temperature. After a certain time, it was found that the rate of the reaction with respect to consumption of $\text{C}_2\text{H}_4(\text{g})$ was $x \text{ mol dm}^{-3} \text{ s}^{-1}$. Which of the following shows the rates of consumption of $\text{O}_2(\text{g})$, formation of $\text{CO}_2(\text{g})$ and formation of $\text{H}_2\text{O}(\text{g})$ respectively, during that time?

	rate / $\text{mol dm}^{-3} \text{ s}^{-1}$		
	$\text{O}_2(\text{g})$	$\text{CO}_2(\text{g})$	$\text{H}_2\text{O}(\text{g})$
(1)	$\frac{3}{x}$	$\frac{2}{x}$	$\frac{2}{x}$
(2)	x	x	x
(3)	$\frac{x}{3}$	$\frac{x}{2}$	$\frac{x}{2}$
(4)	$\frac{1}{x}$	$\frac{1}{x}$	$\frac{1}{x}$
(5)	$3x$	$2x$	$2x$

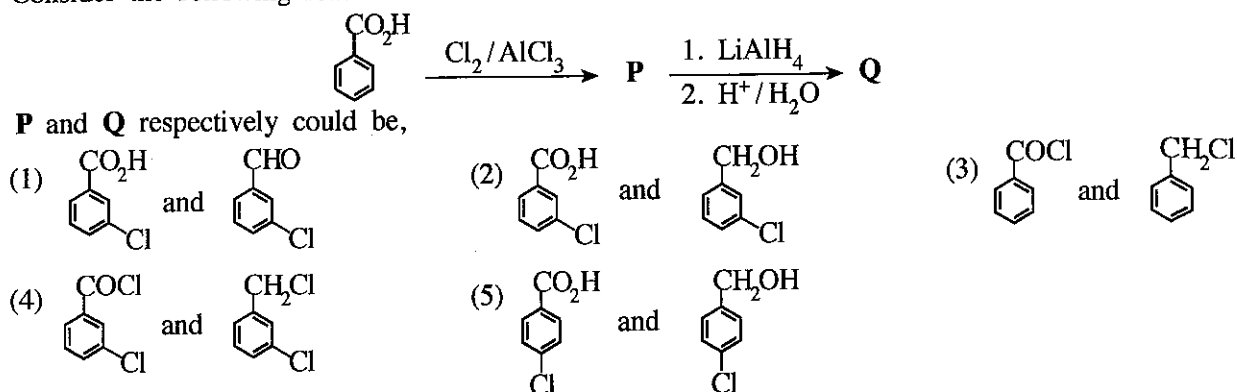
29. Consider the following reaction occurring in a rigid-closed container at temperature T .



The rate of reaction doubled when the concentration of M was doubled. The rate of reaction is $5.00 \times 10^{-4} \text{ mol dm}^{-3} \text{ s}^{-1}$ when the concentrations of M and Q are $1.0 \times 10^{-5} \text{ mol dm}^{-3}$ and 2.0 mol dm^{-3} respectively. The rate constant of the reaction under these conditions is,

- (1) $2.5 \times 10^{-4} \text{ s}^{-1}$ (2) 12.5 s^{-1} (3) 25 s^{-1} (4) 50 s^{-1} (5) 500 s^{-1}

30. Consider the following reaction scheme.



● For each of the questions 31 to 40, one or more responses out of the four responses (a), (b), (c) and (d) given is/are correct. Select the correct response/responses. In accordance with the instructions given on your answer sheet, mark

- (1) if only (a) and (b) are correct.
 (2) if only (b) and (c) are correct.
 (3) if only (c) and (d) are correct.
 (4) if only (d) and (a) are correct.
 (5) if any other number or combination of responses is correct.

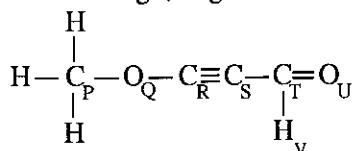
Summary of above Instructions

(1)	(2)	(3)	(4)	(5)
Only (a) and (b) are correct	Only (b) and (c) are correct	Only (c) and (d) are correct	Only (d) and (a) are correct	Any other number or combination of responses is correct

31. Which of the following statement/s is/are correct with regard to 3d-block elements and their compounds?

- (a) Among the 3d-block elements, Sc has the highest density.
 (b) The radii of atoms (Sc to Cu) increase from left to right.
 (c) $[\text{Ni}(\text{NH}_3)_6]^{2+}$ is blue in colour whereas $[\text{Zn}(\text{NH}_3)_4]^{2+}$ is colourless.
 (d) The IUPAC name of K_2NiCl_4 is dipotassium tetrachloronickelate(II).

32. Which statement/s is/are correct regarding the following molecule?



- (a) Atoms labelled P, Q, R and S lie on a straight line.
 (b) Atoms labelled Q, R, S and T lie on a straight line.
 (c) Atoms labelled R, S, T, U and V lie on the same plane.
 (d) Atoms labelled R, S, T and U lie on a straight line.

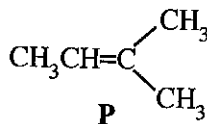
33. Consider the equilibrium $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ at 500 K.

Which of the following statement/s is/are correct for the above equilibrium?

- (a) Expression for the equilibrium constant for the equilibrium is, $K_c = \frac{[\text{NH}_3(\text{g})]^2}{[\text{N}_2(\text{g})]^{\frac{1}{2}}[\text{H}_2(\text{g})]^{\frac{3}{2}}}$
- (b) Expression for the equilibrium constant for the equilibrium is, $K_c = \frac{[\text{N}_2(\text{g})]^{\frac{1}{2}}[\text{H}_2(\text{g})]^{\frac{3}{2}}}{[\text{NH}_3(\text{g})]^2}$
- (c) Expression for the equilibrium constant for the backward reaction is, $K'_c = \frac{[\text{NH}_3(\text{g})]^2}{[\text{N}_2(\text{g})][\text{H}_2(\text{g})]^3}$
- (d) Expression for the equilibrium constant for the backward reaction is, $K'_c = \frac{[\text{N}_2(\text{g})][\text{H}_2(\text{g})]^3}{[\text{NH}_3(\text{g})]^2}$

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34. Which of the following statement/s regarding the reaction between compound **P** and HCl to form an alkyl halide is/are correct?



- (a) The major product is 2-chloro-2-methylbutane.
 (b) A secondary carbocation is formed as an intermediate in this reaction.
 (c) In one of the steps of the reaction, the HCl bond is cleaved to give a chlorine radical (Cl[•]).
 (d) In one of the steps of the reaction, a nucleophile reacts with a carbocation.
35. A binary liquid mixture prepared by mixing two liquids in a closed evacuated container at a given temperature shows a negative deviation from Raoult's Law. Which of the following statement/s is/are correct for this system?
- (a) Total vapour pressure of the mixture is less than the expected total vapour pressure should it behave as an ideal mixture.
 (b) Heat is released when the mixture is formed.
 (c) Number of molecules in the vapour phase of the mixture is greater than the expected number of molecules should it behave as an ideal mixture.
 (d) Heat is absorbed when the mixture is formed.
36. Which of the following statement/s is/are correct with regard to effects on natural cycles (e.g. N₂ cycle) present in our environment due to human activities?
- (a) Fossil fuel burning contributes to elevated levels of CO₂ in the atmosphere.
 (b) CO₂ emission due to respiration of organisms causes elevated levels of CO₂ in the atmosphere.
 (c) Industrial fixation of nitrogen causes increased amounts of nitrogen containing compounds in water and soil.
 (d) Deforestation does not contribute to elevated levels of CO₂ in the atmosphere.
37. Which of the following statement/s is/are correct with regard to halogens, noble gases and their compounds?
- (a) Hypochlorous ion disproportionates rapidly in acidic solutions.
 (b) Xe forms a series of compounds with F₂ gas, among which XeF₄ has a square planar geometry.
 (c) Among the hydrogen halides, HF has the highest bond dissociation energy per mole.
 (d) Boiling points of halogens increase down the group as a result of increasing strength of London forces.
38. Which of the following statement/s is/are correct regarding the Daniell cell when it operates at room temperature? ($E_{cell}^{\circ} = +1.10 \text{ V}$)
- (a) Net electron flow occurs from Zn to Cu.
 (b) The equilibrium $\text{Zn}^{2+}(\text{aq}) + 2\text{e} \rightleftharpoons \text{Zn}(\text{s})$ shifts to the right.
 (c) A liquid-junction potential is created due to the presence of a salt-bridge.
 (d) The equilibrium $\text{Cu}^{2+}(\text{aq}) + 2\text{e} \rightleftharpoons \text{Cu}(\text{s})$ shifts to the right.
39. Which of the following statement/s is/are correct for ideal gases and real gases at constant temperature?
- (a) At very high pressures, the volume of a real gas is higher than that of an ideal gas.
 (b) At high pressures, real gases tend to behave as ideal gases.
 (c) At very high pressures, the volume of a real gas is lower than that of an ideal gas.
 (d) At low pressures, real gases tend to behave as ideal gases.
40. Which of the following statement/s is/are correct regarding some industrial processes?
- (a) The first two steps involved in the manufacture of Na₂CO₃ by Solvay Process are endothermic.
 (b) Production of urea is carried out by a two step process using ammonia and carbon dioxide as the only raw materials.
 (c) The first step involved in the manufacture of nitric acid by Ostwald method is the oxidation of NH₃ gas using O₂ in air in the presence of a catalyst to give NO₂ gas.
 (d) High temperature and low pressure conditions are employed in the manufacture of NH₃ gas using Haber-Bosh process.

- In question Nos. 41 to 50, two statements are given in respect of each question. From the Table given below, select the response, out of the responses (1), (2), (3), (4) and (5), that best fits the two statements and mark appropriately on your answer sheet.

Response	First Statement	Second Statement
(1)	True	True, and correctly explains the first statement
(2)	True	True, but does not explain the first statement correctly
(3)	True	False
(4)	False	True
(5)	False	False

	First Statement	Second statement
41.	Among the oxides of Cr and Mn, CrO and MnO are acidic, while CrO ₃ and Mn ₂ O ₇ are basic.	The acidic/basic nature of the oxides of Cr and Mn is dependant on the oxidation number of the metal.
42.	An acidic buffer solution can be prepared by mixing a weak acid HA(aq) with its sodium salt NaA(aq).	When OH ⁻ (aq) or H ⁺ (aq) ions are added to a buffer solution, the added amounts of OH ⁻ (aq) or H ⁺ (aq) ions are removed through the reactions; OH ⁻ (aq) + HA(aq) → A ⁻ (aq) + H ₂ O(l) and H ⁺ (aq) + A ⁻ (aq) → HA(aq) respectively.
43.	Polyvinyl chloride (PVC) shows some resistance to catching fire.	Polyvinyl chloride (PVC) is a thermosetting polymer.
44.	At a given temperature and pressure the molar volumes of two different ideal gases are different from each other.	At 0 °C temperature and 1 atm pressure, the molar volume of an ideal gas is 22.4 dm ³ mol ⁻¹ .
45.	All compounds having a C=C bond show diastereoisomerism.	Any two isomers which are not mirror images of each other are diastereoisomers.
46.	Hydrogenation of benzene is more difficult than hydrogenation of alkenes.	Addition of hydrogen to benzene results in the loss of aromatic stabilization.
47.	The reaction that takes place between SO ₃ gas and water in the production of sulphuric acid is endothermic.	SO ₃ gas reacts with concentrated H ₂ SO ₄ to give oleum.
48.	Reaction between ammonia and an alkylhalide gives a mixture of primary, secondary and tertiary amines and a quaternary ammonium salt.	Primary, secondary and tertiary amines can react as nucleophiles.
49.	If P + Q → R is a first order reaction with respect to the reactant P, the graph of rate against concentration of P gives a straight line passing through the origin.	Initial rate of a first order reaction is independent of the concentration of reactant(s).
50.	On a sunny day, strong photochemical smog can be seen in a city with heavy traffic congestion.	Photochemical smog is caused entirely by scattering of solar radiation by small particles and water droplets that are emitted by vehicle exhaust systems.

The Periodic Table

1	1 H																		2 He
2	3 Li	4 Be										5 B	6 C	7 N	8 O	9 F	10 Ne		
3	11 Na	12 Mg										13 Al	14 Si	15 P	16 S	17 Cl	18 Ar		
4	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	
5	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe	
6	55 Cs	56 Ba	La- Lu	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn	
7	87 Fr	88 Ra	Ac- Lr	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og	

57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu
89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

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